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Answers

## **Chapter 16 -**

### **Solutions - 16.2**

#### **Concentrations of Solutions ...**

Concentrations of  
Solutions Practice.

Calculate the  
concentration in e. ach  
of the following  
solutions: 10 moles of  
potassium hydroxide in  
5.16 L of solution. 0.5

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moles of calcium chloride in 485 mL of solution. 0.079 moles of magnesium sulfate in 25 mL of solution.

Find the number of moles of solute in each of the following solutions: 2L ...

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SOLUTION  
CONCENTRATIONS**

Concentrations of  
Solutions Date \_\_\_\_\_  
Complete the following  
problems on a

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Worksheet

Answers

separate sheet of paper. Use significant figures. Note: The density of water is 1 g/mL. 1. What is the molarity of a solution that contains 10.0 grams of Silver Nitrate that has been dissolved in 750 mL of water? 10.0!!!!!"! " ...

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How does a solution  
become  
supersaturated? \_\_\_\_\_

SECTION 16.2

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SOLUTIONS (pages 480-486) This section explains how to solve problems involving molarity of a solution, how to prepare dilute solutions from more concentrated solutions, and what is meant by percent by volume and percent by mass.

Molarity (pages 480-482) 1.

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Calculations for Solutions Worksheet and Key 1) 23.5g of NaCl is dissolved in enough water to make 683L of solution . a) What is the molarity (M) of the solution? b) How ...

### **Calculations for Solutions Worksheet and Key**

The concentrations of very dilute solutions are often expressed in parts per million (ppm),

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which is grams of solute per 10<sup>6</sup> g of solution, or in parts per billion (ppb), which is grams of solute per 10<sup>9</sup> g of solution. For aqueous solutions at 20°C, 1 ppm corresponds to 1 µg per milliliter, and 1 ppb corresponds to 1 ng per milliliter.

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Concentrations

Worksheet (Section

12.3) Name \_\_\_\_\_

Period: Measuring

Concentration: There

are several different

ways to measure and

express the

concentration of a

solution. Molarity

(Section 12.3) the term

we learned earlier,

refers to the

concentration of a

solution expressed in

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liter of solution. ...

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